

1. Partial pressure of a gas in a mixture is the product of _____
total pressure and mole fraction
2. Calculate the volume occupied by 0.07 Kg of carbon (IV) oxide gas at a temperature of 303K and a pressure of 30500 Pa assuming ideal behaviour
131.3dm³
3. The relationship between volume and the pressure of a given mass of gas at a given temperature is known as _____
Boyles Law
4. Assuming a 4.00 litre of a sample gas at 1.00 atm compressed to 0.800 litre at constant temperature. Calculate the final pressure of the gas.
50 atm
5. 3 atm expressed in pascal is approximately _____
300000
6. Calculate the volume (in cm³) that 4.00 cm³ of gas at 30 °C will occupy at 130°C when the pressure is kept constant
5.32
7. Gaseous molecules are separated by _____ forces
Intermolecular forces
8. The following properties can describe the state of an ideal gas EXCEPT
None of the above
9. 1 atm is equivalent to _____
760 mm Hg
10. Which of these is NOT true about ideal gases
The interaction between molecules are very strong

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