

CHM205

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1. The mineral form of phosphorous  $\text{Ca}_3(\text{PO}_4)_2$  is also present in

Blood

Flesh

--->> Teeth

Marrow

2. The order of electron affinity in Group 7 elements is

$\text{Cl} > \text{Br} > \text{F} > \text{I}$

$\text{Cl} > \text{Br} > \text{I} > \text{F}$

--->>  $\text{F} > \text{Cl} > \text{Br} > \text{I}$

$\text{F} > \text{Cl} > \text{I} > \text{Br}$

3. The process through which Nitric acid is manufactured is called \_\_\_\_\_

Harber's process

Le-Chatelier's process

Contact process

--->> Ostwald's process

4. In Group 7, the electronic configuration of seven electrons in the outermost shells is

$np^5 nd^2$

$ns^1 np^6$

$np^6 nd^1$

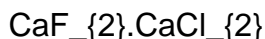
--->>  $ns^2 np^5$

5. Fluorspar is a mineral form of fluorine. Its chemical formula is \_\_\_\_\_

--->>  $\text{CaF}_2$

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6. Nitrogen is quite inert at room temperature because of its

longer N = N bond length

stronger N $\equiv$ N bond

weaker N = N bond

--->> shorter N = N bond length

7. Halogens do not occur free in nature due to their \_\_\_\_\_

Low electronegativity

Zero electronegativity

--->> High reactivity

Low reactivity

8. The compound of arsenic that find extensive use as herbicide is \_\_\_\_\_

--->> Monosodium methyl arsenate

Sodium arsenite

sodium tetraoxo arsenate (V)

Arsenic pentoxide

9. \_\_\_\_\_ may be regarded as the removal of electrons

Reduction

--->> Oxidation

Redox

Electron affinity

10. Ammonifying bacteria in the soil do act on dead plants and animals to produce

Ammonium chloride and ammonium sulphate

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Aqueous ammonia and nitrogen (II) oxide gas

Nitric acid and nitrous acid

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