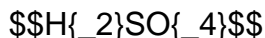
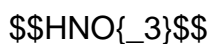


CHM203

=====

1. Which of these is a weak acid?



2. Why are bronsted acid call protic acid?

Because they accept proton

Because they donate protons to water

Because they donate protons to base

---> Because they react by the transfer of proton

3. Give reason to why ethanoate ions are called conjugate base of an ethanoic acid?

---> Because ethanoate can accept proton to become ethanoic acid

Because ethaoate ion can donate another proton to water

Because ethanoate is the ionic form of ethanoic

Because ethanoate is formed by the loss of proton from ethanoic acid

4. Which of these statements is not true about the strength of a conjugate base?

The stronger the acid the weaker will be its conjugated base

---> The stronger the acid, stronger will be its conjugate base

Stronger the acid the higher its PK_a value

Stronger the acid the lower its PK_a value

5. Define acid in bronsted and Lowry perspective.

An acid is a proton acceptor

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An acid is a hydroxyl ion acceptor

An acid is a substance which dissociate to yield proton

6. The pair of acid and its conjugate base or base and its conjugate acid is called?

Conjugate base-proton pair

--->> Conjugate acid-base pair

Conjugate acid pair

Conjugate base pair

7. Hydroxyl ion is a conjugate base of water. True or false?

--->> True

True

partially true

indifferent

8. The strength of an acid is measured by its ability to ____?

Be corrosive

--->> Dissociate almost completely in aqueous solution

React with base to form salt and water

Dissociate completely into hydrogen ions and hydroxyl ions

9. Acids are electron pair acceptors and bases are electron donors according to _____?

Bronsted-lowry

Lowry

--->> Lewis

bronsted

10. Define acid according to Arrhenius; 1884.

--->> An acid is a substance which ionizes in aqueous solution to produce

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An acid is a proton donor

An acid a substance which ionizes an aqueous solution to produce hydroxyl ions

An acid is a proton acceptor

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