

CHM103

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1. The pressure exerted when an equilibrium is established between the number of particles evaporating and condensing within a given system is known as

Standard pressure

--->> vapour pressure

Total pressure

partial pressure

2. Under what conditions does the behaviour of real gases deviate from that predicted by the ideal gas law?

High Pressure, High Temperature

Low Pressure, low Temperature

low Pressure, High Temperature

--->> High Pressure, Low Temperature

3. One of these form a crystal with a bcc structure

--->> CsCl

NaCl

Ni

S

4. The point whereby the temperature in which the prevailing vapour pressure of the liquid is equal to the ambient atmospheric pressure is called?

melting point

--->> boiling point

reference point

standard point

5. Which of these statements is true of mixtures?

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--->> Very little or no exchange of energy occurs during the formation of mixtures

A mixture cannot be easily separated into its constituents

A mixtures can only be separated into its constituents by chemical methods

6. Helium molecule has a mass of 4.0026 g and is moving at a speed of 30 m/s. The classical kinetic energy of the gas is?

--->> 1.80 J

18.0 J

0.180 J

0.0180 J

7. According to Clausius-Clayperon equation____

vapour pressure of any substance increase linearly with temperature

--->> vapour pressure of any substance does not increase linearly with temperature

vapour pressure of any substance increase inversely with temperature

vapour pressure of any substance does not increase inversely with temperature

8. A plot of V against T as linear with zero intercept and slope, equal to K is graphical representation of _____

Boyle's law

Avogadro's law

--->> Charles' law

Gay Lussac's law

9. Which of the following is a pure substance made up of one type of atom?

A mixture

Compound

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--->> Element

10. When 0.25 L of liquid nitrogen ($d = 0.807 \text{ g/mL}$) is vaporized, what volume does the resulting gas occupy at $25 \text{ }^\circ\text{C}$ and 5.00 atm

--->> 35L

54L

17L

93L

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