
percent yield of product

[CHM291] The total mass of each element in a compound depends on the number of
its _____
atoms

[CHM291] the determination of the ratio of the individual elements in a known mass of
the compound is also known as _____
determination of empirical formula

[CHM291] _____ is prepared from an aqueous solution of ethylenediamine and virtually
any cobalt(II) salt
Ethylenediamine cobalt (III) chloride

[CHM291] the simplest way to express information about the atoms that constitute any
given chemical compound is _____
chemical formula

[CHM291] _____ is composed of atoms of two or more elements chemically
combined in definite proportions
chemical compound

[CHM291] What is the product of the reaction between one mole of barium ion, from
solution of $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ and with 1 mole of sulfate ion, from solution of 1 mole of
 Na_2SO_4
1 mole of barium sulfate precipitate

[CHM291] only a limited amount of product forms from given amounts of reactants
because _____
chemicals react stoichiometrically

[CHM291] The empirical formula of a strip of aluminum weighing 0.690g is ignited
yielding an oxide that weighs 1.300g
 $\text{\$Al}_2\text{O}_{3}\text{\$}$

[CHM291] Hexaamminecobalt(III) salts may be prepared by any of these three
methods which depends the on oxidation of cobalt(II) ion in ammoniacal solution except

hydrated potassium aluminum sulfate

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