

0 to 4

[CHM204] Energy released when gaseous ions are converted into the solid ionic compound is
crystal lattice energy

[CHM204] The strength of "ionic bonds" in compounds is estimated by _____
Born-Haber cycles

[CHM204] Electronegativity is a properties that _____ across the period
increases

[CHM204] The element with the lowest electron affinity with value of 0.7 is _____
caesium

[CHM204] What type of energy is required for the equation: $\text{Hg}_{(l)} \rightarrow \text{Hg}_{(g)}$?
heat of vaporization

[CHM204] electron affinities become more _____ as we move from left to right
across a period
exothermic

[CHM204] Which law states that the energies of the reactions in the first sequence must add up to the heat of formation
Hess's law

[CHM204] What type of energy is required for the equation: $\text{H}_{2(g)} \rightarrow 2\text{H}_{(g)}$?
bond dissociation energy

[CHM204] The ionization energies _____ across the period
increase

Whatsapp: 08089722160 or click here for TMA assistance

Practice E-exams & Chat with course mates on [noungeeks.net](#)