

used with the____
Thiopsulphate

[CHM191] The most commonly used external indicator for iodine titration is_____
starch solution

[CHM191] The _____is the amount of heat released or observed for a given
amount of reactants or products
heat of reaction

[CHM191] The study of thermal changes in chemical and physical processes is known
as_____
Thermodynamics

[CHM191] For this reason and in order to avoid complications in the end point detection,
the starch indicator is added at the_____ colouration of the solution which is near the
end point
Light yellow

[CHM191] Thus using starch as indicator, the colour variations of the solution depend
on _____formation of the starch with iodide ion formed in the last stage of the titration.
Complex

[CHM191] In the _____, standard solutions of iodine are used to estimate directly the
concentrations of some oxidizable species.
Direct method

[CHM191] The starch solution forms a _____complex with the tri-iodide ion during
the titration that is rapidly discharge at the end point
Blue black

[CHM191] Complete the colour change in iodometry analysis. _____(initially)
Red-brown

[CHM191] The oxidation potentials of Na is_____
2.71

Whatsapp: 08089722160 or click here for TMA assistance

Practice E-exams & Chat with course mates on [noungeeks.net](#)