

mole (molar mass = 12 g.mol) deposited is ____
10

[CHM409] Which of the following manufacturing process depends on electrolysis?
Down

[CHM409] If the free energy change of a cell is -457 kJ/mol, what is the electrode potential of the cell ($n = 2$)
+2.37 J

[CHM409] A pH meter is an example of ----- electrode
Ion selective electrode

[CHM409] The temperature coefficient of a half wave potential often lies between -2 and ____
2

[CHM409] Irreversible thermodynamic cell reaction attained equilibrium ____
Slowly

[CHM409] Over potential is positive for ____ polarization
Anodic

[CHM409] A reversible cell reaction attained thermodynamic equilibrium ____
Rapidly

[CHM409] When charging a lead acid battery, lead sulphate at the anode is ----- to Pb
Reduced

[CHM409] The quantity of charge required to obtain one mole of sodium from NaOH is
1F

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