

likely a\_\_\_\_  
Unimolecular

[CHM407] The molecularity of a reaction cannot be one of the following\_\_\_\_  
4

[CHM407] The number of molecules associated with the rate determining step is  
called\_\_\_\_  
Molecularity

[CHM407] A slope of a plot of  $\ln(\text{rate constant})$  versus  $1/\text{Temperature}$  gave 0.5.  
Calculate the activation energy of the reaction.  
0.5 J/mol

[CHM407] Reactions initiated by free radicals is called\_\_\_\_  
Chain reaction

[CHM407] A complex reaction in which the reactant gives more than one product is  
called \_\_\_\_  
Parallell reaction

[CHM407] Given that the rate constants for a first order reaction at 25 and 60  $^{\circ}\text{C}$   
are  $2 \times 10^{-3}$  and  $8 \times 10^{-3}$  s/mol. Calculate the activation energy for the reaction.  
32.33 kJ/mol

[CHM407] One of the following reaction does not have complicated mechanism\_\_\_\_  
Esterification

[CHM407] A second order reaction made to behave as a first order is called\_\_\_\_  
pseudo first order

[CHM407] One of the following parameters cannot be obtained from a linear plot of  $\log$   
(rate/Temperature) versus  $1/\text{Temperature}$   
Activation energy

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